Rotationally Resolved Spectra of Isovalent NbCr and VCr†

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Resonant two-photon ionization spectroscopy has been used to study the isovalent metal molecules NbCr and VCr. The first experimental observation of $^{93}$Nb$^{52}$Cr yielded a ground-state rotational constant of $B''_0 = 0.141 \pm 0.003 \text{ cm}^{-1}$, which corresponds to a bond length of $r''_0 = 1.8940 \pm 0.0003 \text{ Å}$. The excited state, which lies at 14 440 cm$^{-1}$, has $B'_0 = 0.144 \pm 0.001 \text{ cm}^{-1}$ and $r'_0 = 1.8720 \pm 0.0001 \text{ Å}$. The bond energy of NbCr is measured as 24.409 ± 5 cm$^{-1}$ (3.0263 ± 0.0006 eV) by the onset of predissociation in a congested vibronic spectrum. For $^{51}\text{V}^{52}\text{Cr}$, the ground state is identified as $^2\Delta_{52}$ with $B'_0 = 0.219 \pm 0.001 \text{ cm}^{-1}$ and $r'_0 = 1.7260 \pm 0.0011 \text{ Å}$. The excited state, which lies at 14 371 cm$^{-1}$, has $B'_0 = 0.221 \pm 0.001 \text{ cm}^{-1}$ and $r'_0 = 1.7201 \pm 0.0011 \text{ Å}$. On the basis of these results, a consistent set of multiple bonding radii are presented for the V, Cr, Nb, and Mo atoms, and the bond lengths ($r_n$) of VMo and NbMo are predicted to be 1.859 and 2.013 Å, respectively. A comparison of bond energies for the NbM molecules is also presented, along with an analysis of the trend in bond energies as the metal M is varied.

I. Introduction

Many experimental and theoretical studies have examined the group 5 and group 6 transition metal dimers,1–11 characterizing the homonuclear diatomics $\text{V}_2$, $\text{Cr}_2$, $\text{Nb}_2$, and $\text{Mo}_2$, as well as the intragroup heteronuclear diatomics $\text{VNb}$ and $\text{CrMo}$. In contrast, rather little is known about the mixed group 5/6 dimers. To our knowledge, only five previous studies exist concerning the mixed group 5/6 dimers. A Knudsen effusion mass spectrometer. A complete description of the experimental details has been previously published.19

The greatest concentrations of the heteronuclear molecules VCr and NbCr that could be produced are estimated to be less than 5% of the concentrations of the homonuclear molecules $\text{V}_2$ and $\text{Nb}_2$, respectively. This is presumably because the high bond energies of $\text{V}_2$ (2.75 eV)20 and $\text{Nb}_2$ [5.24(13) eV]21, as compared to VCr (calculated $D_b = 2.262 \text{ eV}$)18 and NbCr (3.026 eV, see below), facilitated the displacement reactions $\text{V} + \text{Cr} \rightarrow \text{V}_2 + \text{Cr}$ and $\text{Nb} + \text{Cr} \rightarrow \text{NbCr} + \text{Cr}$, thereby leading to the removal of the target molecules from the molecular beam. To overcome this problem, the samples were enriched in Cr, as described above. Even so, the heteronuclear molecules ($^{51}\text{V}^{52}\text{Cr}$ and $^{93}\text{Nb}^{52}\text{Cr}$) showed an extremely weak signal strength, even when on resonance, compared to the homonuclear signals for $\text{V}_2$ and $\text{Nb}_2$. Because of this weak signal strength, it was impossible to record spectra for any of the minor isotopes of VCr or NbCr, and it was likewise impossible to resolve the hyperfine structure of NbCr.

II. Experimental Section

Diatomic NbCr and VCr were formed by pulsed-laser ablation (Nd:YAG, 532 nm, 9 mJ per pulse) of alloy disks with mole ratios of 4:1 (Cr:Nb or Cr:V), followed by supersonic expansion in helium carrier gas. After expansion into vacuum, the molecular jet was skimmed, and the collimated supersonic beam was admitted into a second chamber where it was exposed to tunable visible radiation produced by a Nd:YAG pumped dye laser. Photoionization of any excited states so produced was accomplished using the 248-nm radiation produced by an excimer laser operating on a KrF mixture. The ionized molecules were then detected in a reflectron time-of-flight mass spectrometer. A complete description of the experimental details has been previously published.19

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metals, these effects are usually small unless the perturbing states lie very close in energy. For the ground state, which is shown in the next paragraph to be a $^2\Sigma_g^+$ state, the obvious perturber is the higher-energy $^2\Sigma_u^+$ spin–orbit component. For such a situation, the true rotational constant, $B_{\text{true}}$, is related to the effective rotational constant, $B_{\text{eff}}$, by the formula\(^{(3.1)}\)

$$B_{\text{eff}} = B_{\text{true}} + \frac{2B_{\text{true}} \Sigma}{\Delta\Lambda}$$

For the VCr molecule, $A$ is expected to fall between the values of the spin–orbit parameter, $-\xi_{5d}$, for V and Cr, placing $A$ in the range of $-177$ to $-243$ cm\(^{-1}\).\(^{(24)}\) Thus, the value of $B_{\text{true}}$ is slightly modified compared to the fitted value of $B'_{\text{eff}}$, giving $B''_{\text{true}} = 0.2203(28)$ cm\(^{-1}\). This value corresponds to a bond length of $r''_0 = 1.7260(11)$ Å. Here and throughout this article, we report 1σ error limits in units of the last reported digit in parentheses.

The similar intensities of the R and P branches, along with the relative weakness of the Q branch, for which all of the lines overlap, imply that this is a parallel ($\Delta\Omega = 0$) transition. The observation of R(2.5) as the first R line then identifies the band as an $\Omega' = 2.5 \rightarrow \Omega'' = 2.5$ transition. As such, the first P line should be P(3.5), but this line is not readily identified in the spectrum. A candidate line lies slightly to the blue of the P(4.5) line, but it lies too far to the red to be included in the least-squares fit as P(3.5). The P(3.5) line appears to be displaced approximately 0.234 cm\(^{-1}\) from its expected position by a perturbation with another state. Weak, unassigned features near 14 370 cm\(^{-1}\) may be transitions to rotational levels of this perturbing state. Likewise, the R(10.5) and P(12.5) lines, which terminate on the same rotational level of the upper state, are both shifted to lower energy, again indicating a perturbing state lying close in energy. Because of these apparent perturbations, P(3.5), P(12.5), and R(10.5) were omitted from the least-squares fit. Fitted spectroscopic constants are given in Table 1, along with the residuals in the least-squares fit.

An attempt was made to measure the bond energy of VCr by locating a predissociation threshold in a dense manifold of vibronic levels, as has been previously done for a number of diatomic metal molecules.\(^{4,7,20,25,26}\) This attempt was unsuccessful because of the weakness of the VCr signal, combined with the large number of V2 transitions at higher energies. Even under the best conditions for production of VCr, on resonance,

![Figure 1. Rotationally resolved scan over the 14 371 cm\(^{-1}\) band of $^{51}$V$^{52}$Cr. From the analysis of this $\Omega' = \frac{1}{2} \rightarrow \Omega'' = \frac{5}{2}$ band, the ground-state bond length of VCr is deduced to be $r''_0 = 1.7260(11)$ Å.](image)

<table>
<thead>
<tr>
<th>rotational line</th>
<th>observed line position$^a$</th>
</tr>
</thead>
<tbody>
<tr>
<td>P(4.5)</td>
<td>14 369.138 (−27)</td>
</tr>
<tr>
<td>P(5.5)</td>
<td>14 368.692 (−6)</td>
</tr>
<tr>
<td>P(6.5)</td>
<td>14 368.256 (8)</td>
</tr>
<tr>
<td>P(7.5)</td>
<td>14 367.822 (24)</td>
</tr>
<tr>
<td>P(8.5)</td>
<td>14 367.423 (7)</td>
</tr>
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<td>P(9.5)</td>
<td>14 366.998 (20)</td>
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<tr>
<td>P(10.5)</td>
<td>14 366.623 (−15)</td>
</tr>
<tr>
<td>P(11.5)</td>
<td>14 366.183 (20)</td>
</tr>
<tr>
<td>P(12.5)</td>
<td>14 364.427 (−26)</td>
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<tr>
<td>R(2.5)</td>
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</tr>
<tr>
<td>R(3.5)</td>
<td>14 373.101 (−17)</td>
</tr>
<tr>
<td>R(4.5)</td>
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<tr>
<td>R(5.5)</td>
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<tr>
<td>R(6.5)</td>
<td>14 374.478 (−11)</td>
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<tr>
<td>R(7.5)</td>
<td>14 374.914 (20)</td>
</tr>
<tr>
<td>R(8.5)</td>
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<tr>
<td>R(9.5)</td>
<td>14 375.870 (9)</td>
</tr>
<tr>
<td>R(10.5)</td>
<td>14 376.849 (−13)</td>
</tr>
</tbody>
</table>

$^a$ All numerical values are given in wavenumbers (cm\(^{-1}\)). Residuals in the fit are given in parentheses in units of 0.001 cm\(^{-1}\). $^b$ Line positions are determined by calibrating the etalon with the known frequencies of the I\(_2\) absorption spectrum. All line positions are corrected by the Doppler shift experienced by $^{51}$V$^{52}$Cr as it translates toward the radiation source. The errors associated with the observed line positions are obtained by a least-squares fit of the line positions to the formula $r = v_0 + B'J(J + 1) - B''J(J + 1)$. The resulting values of the parameters are $r_0 = 14 371.0648(82)$ cm\(^{-1}\), $B'=0.2199(26)$ cm\(^{-1}\), $B''=0.2213(53)$ cm\(^{-1}\), $r''_0 = 1.7260(11)$ Å, and $r''_0 = 1.7201(11)$ Å, where the 1σ error limits are given in parentheses, in units of the last digit quoted. In the calculation of $r''_0$, the effects of the spin–uncoupling operator, which couples the $^2\Sigma_g^+$ ground state with the $^2\Sigma_u^+$ excited spin–orbit component, have been included, even though this only leads to a minor change in the estimated bond length of 0.0005 Å.

**B. NbCr**. A scan of the excitation laser from 12 000 to 18 000 cm\(^{-1}\) revealed only one strong transition in NbCr, at 14 440 cm\(^{-1}\). The first attempt to rotationally resolve this transition produced a spectrum without clear rotational lines. The presumed cause of the poor resolution was the hyperfine splitting induced by the large nuclear magnetic moment of $^{93}$Nb ($I = \frac{3}{2}$), 6.167 nuclear magnetons.\(^{(27)}\) Recognizing that the hyperfine splitting in a Hund’s case-(a) or case-(c) molecule decreases with increasing $J$ as $1/J$,\(^{(28,29)}\) we decided that impeding the cooling effects of the supersonic expansion and populating higher rotational levels might produce narrower, recognizable lines. To warm the molecular beam, a 3.5-in.-long extender with an exit orifice of $\frac{1}{8}$ in. (0.95 cm) was affixed to the end of the vaporization block. This modification moved the expansion orifice much closer to the skimmer, reduced the carrier gas pressure at the nozzle orifice, and effectively moved the probe region much closer to the expansion orifice, in units of nozzle diameter. All of these factors contributed to a reduction in the degree of supersonic cooling, as compared to the standard exit orifice diameter of 0.1 cm. The rotationally resolved spectrum recorded with this extender is displayed in Figure 2. The hyperfine broadening of the low-$J$ lines is evident, with lines below R(6.5) and P(9.5) disappearing into a sea of noise.

As was found for VCr, this band displays R and P branches of comparable intensity and has a much weaker Q branch, implying that it is another parallel ($\Delta\Omega = 0$) transition. The lack of assignable lines in the Q branch, combined with our inability to follow the P and R branches down to their first lines,
prohibits a definitive experimental assignment of $\Omega''$ and $\Omega'$ at this time. It also makes the absolute calibration of the predissociation threshold was accomplished by collecting niobium atomic transitions simultaneously in the ab initio calculations of Andersson, where all of the molecular orbitals listed are bonding in character. The bond length of $51\text{V}^{2}\text{Cr}$, $r_0 = 1.700\, \text{Å}$ and $r_e = 1.720\, \text{Å}^{18}$. Previous work on CrMo and VNb has shown that the bond lengths of these molecules are also nearly equal to the averages of their constituent homonuclear molecules.

A previous photoelectron spectroscopic study by Alex and Leopold, in combination with a CASSCF/CASPT2 calculation by Andersson, has provided considerable insight into the low-lying states of VCr. The $^3\Delta_g$ ground state is found to have a high vibrational frequency ($\omega_v = 520 \pm 8\, \text{cm}^{-1}$, $\omega_v x_v = 7.2 \pm 0.9\, \text{cm}^{-1}$ for $^{51}\text{V}^{2}\text{Cr}$), indicating that multiple bonding similar to that found in V$_2$ and Cr$_2$ occurs. This is even more apparent in the first excited state observed, which is a $^3\Sigma^+$ state with the $^3\Delta_g$ excited spin–orbit component, have been included, even though this only leads to a minor change in the estimated bond length of 0.0001 Å.

Theoretical predictions ($r_e = 1.700\, \text{Å}^{16}$ and $r_e = 1.720\, \text{Å}^{18}$). Previous work on CrMo and VNb has shown that the bond lengths of these molecules are also nearly equal to the averages of their constituent homonuclear molecules.

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assignment of the upper- and lower-state \( \Sigma^+ \) state derives primarily from d \( \sigma \)-orbital contributors to these mixed orbitals. Thus, the electronic transition we observe probably corresponds to a \( 2 \Delta \) lower state.

Of greater relevance to the \( 2 \Delta \) state is the hyperfine splitting in a \( 2 \Sigma^+ \) state, which was found in VCr, because for a Hund’s case-(b) \( 2 \Sigma^+ \) state, the hyperfine splitting should not decrease as rapidly with increasing \( J \) as is observed in our spectrum. Indeed, if the spin−rotation constant is small, the hyperfine splitting in a \( 2 \Sigma^+ \) state remains nearly constant as \( N \) increases, leading to Hund’s case-(b) hyperfine coupling.

Given that the rotationally resolved band is most likely a \( 2 \Delta_{5/2} \) \( \leftrightarrow \) \( 2 \Delta_{3/2} \) transition analogous to that observed in VCr, the hyperfine energy in each \( 2 \Delta \) state should follow the expected behavior for a Hund’s case-(a) system, for which the hyperfine energy of each state follows the formula

\[
E_{hf}(S, \Lambda, \Sigma, \Omega, I, J, F) = \frac{h\Omega}{2J(J+1)} \left[ F(I+1) - I(I+1) - J(J+1) \right]
\]

to first order in perturbation theory. In this formula, the hyperfine parameter \( h \) is given by

\[
h = a\Lambda + \left( b_f + \frac{2}{3} \right) \Sigma
\]

where

\[
a = 2.000 \beta_{g} \beta_{n} (r^{-3})
\]

\[
b_f = g_f g_b \beta_{b} \frac{\delta \gamma}{3 \psi(0)}^2
\]

and

\[
c = \frac{3}{2} g_{e} g_{f} b_{b} (3 \cos^2 \theta - 1)(r^{-3})
\]

Here, \( g_f \) is the electronic \( g \)-factor, \( g_l \) is the nuclear \( g \)-factor, given by the nuclear magnetic dipole moment in nuclear magnetons divided by the nuclear spin \( I \); \( \beta_c \) is the Bohr magneton; \( \beta_n \) is the nuclear magneton; \( \theta \) is the angle between the internuclear axis and the vector from the magnetic nucleus to the electron; the expectation values provide averages for the single unpaired electron; and \( |\psi(0)|^2 \) provides the probability density for finding the electron at the magnetic nucleus.

The numerical factors \( g_l \beta_b \beta_n \) combine to give the value 0.003 186 cm\(^{-1}\)bohr\(^2\).

Although it has not been possible for us to analyze the hyperfine structure observed in the spectrum of NbCr because of low signal intensity, the magnitude of the observed splittings is not unreasonable. A range of possible values for the hyperfine constant, \( h \), can readily be predicted using atomic hyperfine parameters for \( ^{93}\text{Nb} \). The maximum possible value of the Fermi contact parameter, \( b_f \), can be estimated from the \( A_{iso} \) value listed spectroscopic studies at 4485 ± 15 cm\(^{-1}\). From a \( 2 \Sigma^+ \) state, the only electronic transitions allowed in Hund’s case (a) are to another \( \Sigma^+ \) state or to a \( 2 \Pi \) state. The observed band shows none of the characteristics of these transitions, as a \( \Sigma^+ \) \( \rightarrow \) \( \Sigma^+ \) transition would exhibit only R and P branches, corresponding to \( \Delta N = \pm 1 \), and a \( 3 \Pi \) \( \rightarrow \) \( 2 \Sigma^+ \) transition would display four apparent branches for each \( 3 \Pi_{\pm 1} \rightarrow \Sigma^+ \) sub-band. Our band clearly displays three branches and, therefore, does not originate from a \( 2 \Sigma^+ \) lower state. Although there is no absolute experimental proof that the observed band corresponds to a \( 2 \Delta_{5/2} \) \( \leftrightarrow \) \( 2 \Delta_{3/2} \) transition, this appears to be the only reasonable alternative.

Our inability to analyze the low-\( J \) portion of the spectrum results from the large hyperfine splitting found for the low-\( J \) lines. This, too, argues against a \( 2 \Sigma^+ \) ground state, because for a Hund’s case-(b) \( 2 \Sigma^+ \) state, the hyperfine splitting should not decrease as rapidly with increasing \( J \) as is observed in our spectrum. Indeed, if the spin−rotation constant is small, the hyperfine splitting in a \( 2 \Sigma^+ \) state remains nearly constant as \( N \) increases, leading to Hund’s case-(b) hyperfine coupling.
TABLE 3: Multiple Bonding Radii for V, Cr, Nb, and Mo

<table>
<thead>
<tr>
<th>molecule</th>
<th>measured ( r_0 (\text{Å}) )</th>
<th>reference</th>
<th>predicted ( r_0 (\text{Å}) )</th>
<th>difference (Å)</th>
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<td>V(_2)</td>
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<td>1.779</td>
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<tr>
<td>V Nb</td>
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<td>10</td>
<td>1.934</td>
<td>0.012</td>
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<tr>
<td>Nb(_2)</td>
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<td>5</td>
<td>2.088</td>
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<td>V Cr</td>
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<td>1.940</td>
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<td>1.944</td>
<td>-0.004</td>
</tr>
</tbody>
</table>

a The multiple bonding radii were determined by a least-squares fit using the measured bond lengths of the molecules listed here. b The predicted values were determined by adding the multiple bonding radii of the constituent atoms.

in Table B1 of Weltner’s compilation, giving \( b_H \leq 0.2198 \text{ cm}^{-1} \). This maximum value could result if the unpaired electron were in an orbital that had purely 5s\(_{\text{NB}}\) character, a possibility that is inconsistent with the proposed \( d\sigma^2d\sigma^2\sigma^2\)\(_{\text{A}}\) ground state. On the other hand, the expected values of \( a \) and \( c \) can also be deduced from the dipolar constants listed in Table B1 of Weltner’s book. For the proposed \( \Delta\sigma^2 \) ground state, the expected values are \( a = 0.015 \text{ cm}^{-1} \) and \( c = -0.013 \text{ cm}^{-1} \), giving \( h\Omega = 0.065 \text{ cm}^{-1} \). This value is consistent with the observed hyperfine widths, although simulations of the spectrum require that the upper state must then have a value of \( h\Omega \approx 0.27 \text{ cm}^{-1} \). This result suggests that the upper state may have some \( d\sigma^2d\sigma^2\sigma^2\sigma^2\)\(_{\text{A}}\) character, allowing a significant Fermi contact interaction \( (b_H) \) to develop.

To our knowledge, no previous work, either theoretical or experimental, exists on the NbCr diatomic molecule. Nevertheless, useful comparisons to related molecules can be made. The bond length that was found for NbCr (1.8940 Å) is again remarkably close to the average (1.8826 Å) of the homonuclear molecules Nb\(_2\) (\( r_0 = 2.0793 \text{ Å} \)) and Cr\(_2\) (\( r_0 = 1.6858 \text{ Å} \)). This result is similar to previous findings comparing the bond lengths of Cr\(_2\), Mo\(_2\), and CrMo and of V\(_2\), Nb\(_2\), and V Nb. In fact, if one adopts the multiple bonding radii values of \( r(V) = 0.8896 \text{ Å} \), \( r(Cr) = 0.8443 \text{ Å} \), \( r(Nb) = 1.0441 \text{ Å} \), and \( r(Mo) = 0.9718 \text{ Å} \), the \( r_0 \) bond lengths of V\(_2\), Cr\(_2\), Nb\(_2\), Mo\(_2\), V Cr, V Mo, and Cr Mo are predicted to an accuracy of ±0.012 Å as the sum of the multiple bonding radii, as shown in Table 3. This demonstrates that the chemical bond is similar in all of these species and allows the bond lengths of the as-yet-unobserved molecules V Mo and Nb Mo to be predicted as \( r_0(\text{V Mo}) = 1.861 \text{ Å} \) and \( r_0(\text{Nb Mo}) = 2.016 \text{ Å} \).

C. Bond Energies of the NbM and VM Diatomic Metal Compounds. With the measurement of \( D_0(\text{Nb V}) = 3.0263(6) \text{ eV} \), the bond energies of the diatomic compounds of niobium with the 3d transition metals Ti,\(^{20}\) V,\(^{7}\) Cr,\(^{25}\) Co,\(^{25}\) and Ni\(^{25}\) are now accurately known. These energies are listed in Table 4, along with other information about the known or surmised ground states of these molecules. These values allow us to examine the trends in the chemical bonding of the 3d metals with niobium. At first glance, it would appear that the bond energies are fairly uniform, falling within the range of 2.7–3.1 eV for Ti Nb, Nb Cr, Nb Co, and Nb Ni. However, V Nb severely breaks this uniformity in bond energy with \( D_0(\text{V Nb}) = 3.789 \text{ eV} \).

Efforts to explain the variation in bond energies among a series of transition metal molecules have, in the past, focused on the concept of promotion energy, which can be loosely defined as the amount of energy required to prepare the transition metal atom for bonding. This concept works reasonably well for the transition metal hydride cations, MH\(^+\), in which only a single bond can be formed.\(^{38}\) For such molecules, the promotion energy can be defined as the amount of energy required to excite the atom from its ground spin–orbit level to a hypothetical state in which the spin of the s\(^1\) electron that will form the bond is decoupled from that of the remaining d\(_s\) electrons.\(^{38,39}\) The energy of this spin-decoupled state is taken as the average of the spin–orbit averaged energies of the high-spin and low-spin couplings of the d\(_s\)\(^1\) states of the atom. By adding the promotion energy to the measured bond energies, one obtains the bond energy that would be expected if the ground-state atoms were perfectly set up for the formation of an \( \sigma \) bond, such that no promotion energy would be required.

Adapting this concept to the transition metal dimers, we would expect that in cases in which the \( \sigma^* \) orbital is empty in the diatomic ground state, both atoms would require promotion to the spin-uncoupled d\(_s\)\(^1\) states prior to bond formation. The other separated-atom limit that can lead to an \( \sigma^* \) bond is the d\(_s^0\) + d\(_s^2\) asymptote, which is usually much higher in energy for neutral transition metal atoms. Accordingly, we will not consider it further here. On the other hand, for molecules in which the \( \sigma^* \) orbital is singly occupied, only one atom requires promotion prior to bond formation. In the former case, the intrinsic bond energy is given by \( D_0(\text{AB}) + P(A) + P(B) \), whereas for the latter, the intrinsic bond energy is given by \( D_0(\text{AB}) + P(A, B) \). Here, the promotion energies of atoms A and B are given by \( P(A) \) and \( P(B) \), and the lesser of these two quantities is given by \( P(A, B) \). The relevant promotion energies, \( P(A) \) and \( P(B) \), are listed in Table 4 for the NbM and V M molecules, along with the intrinsic bond energies \( D_0(\text{AB}) + P(A) + P(B) \) or \( D_0(\text{AB}) + P(A, B) \).

To apply these ideas to the transition metal dimers, we must know how many s-based electrons are present in the molecule. For homonuclear molecules, it is not difficult to count the number of s electrons involved in the bond, at least for examples such as \( V_2 \), which has a \( \left[ ... \sigma_4^g \sigma_5^g \sigma_3^g \sigma_1^d \sigma_2^d \sigma_3^d \right] \), \( \text{g}^4 \) ground state.\(^{27,33}\) In this case, no matter how the \( 4d \) and \( 5s \) orbitals are hybridized, there must be a net total of two \( \sigma \) and two \( \sigma^* \) electrons, implying that the ground state derives from a 3d\(_s^4\)\(^1\) + 3d\(_s^4\)\(^1\) separated-atom limit. The situation becomes more problematic in the VNb molecule, where the ground \( \Sigma^- \) state derives from the \( \left[ ... \right] \text{14}^2 \text{15}^a \text{7}^a \text{7}^a \text{2}^d \text{2}^d \text{2}^d \) configuration.\(^7\) In this case, the lack of inversion symmetry removes the restriction that double occupation of the 14\( \sigma \) and 15\( \sigma \) orbitals must give a net total of two \( \sigma \) and two \( \sigma^* \) electrons. Nevertheless, it is probably a good approximation to assume that the 14\( \sigma \) and 15\( \sigma \) electrons in V Nb correspond to two \( \sigma \) and two \( \sigma^* \) electrons. This is because of the similar ionization energies of vanadium (IE = 6.746 eV)\(^{36}\) and niobium (IE = 6.759 eV),\(^{40}\) which imply nearly equal sharing of electrons in the V Nb molecule. Thus, the ground state of V Nb correlates to a 3d\(_s^4\)\(^1\) + 3d\(_s^4\)\(^1\) \( \Sigma^- \) separated-atom limit in which both vanadium and niobium are promoted. More problematic still are the molecules that have an odd number of valence \( \sigma \) electrons, such as Ti Nb and Nb Ni, which have ground states of \( \left[ ... \right] \text{14}^2 \text{15}^a \text{7}^a \text{7}^a \text{2}^d \text{2}^d \text{2}^d \) \( \Sigma^- \) and \( \left[ ... \right] \text{14}^2 \text{15}^a \text{7}^a \text{7}^a \text{2}^d \text{2}^d \text{2}^d \) \( \Sigma^+ \) respectively. In these cases, the extent of \( s-d \) hybridization strongly affects the net number of \( \sigma \) and \( \sigma^* \) electrons that are assigned to the molecule. Finally, an additional significant problem in using promotion energies
to rationalize the measured bond energies of the diatomic transition metal compounds is that, in many cases, the ground electronic state is unknown.

Having noted these caveats, it nevertheless makes sense in the NbM molecules, where M is a 3d metal, to consider the 14σ orbital to be mainly of $5s_{Nb} + 4s_{M}$ bonding character and the 15σ orbital to be mainly of $4d_{Nb} + 3d_{M}$ bonding character, at least in the early transition metal dimers such as TiNb. In combinations involving the late transition metals, such as NbNi, the compositions of these two σ orbitals can reverse, because of the drop in energy of the d-based orbitals as the nuclear charge of M is increased. Nevertheless, the 16σ orbital, which is nominally antibonding, can be considered to be primarily of σ character because, in molecules of this sort, this orbital is so diffuse that it is more appropriately considered nonbonding. This lowers its energy compared to the $3d_{orb}$ orbital, which is strongly antibonding. With these thoughts in mind, it makes sense to consider the TiNb, VNb, and NbCr molecules as correlating to a $d_{s1} + d_{s2}$ separated-atom limit, whereas the NbCo and NbNi molecules, in which the 16σ orbital is presumed or known to be singly occupied, presumably correlate to a $d_{s1} + d_{s2}$ separated-atom limit.

The intrinsic bond energies adopted in Table 4 assume that both transition metal atoms are promoted to $d_{s1} + d_{s2}$ separated-atom limits for the cases of TiNb, VNb, NbCr, TiV, and V2, whereas only a single atom is promoted in the cases of NbCo, NbNi, and VNi. The resulting values are given in bold type under the heading of $D_{0}(AB) + P(A) + P(B)$ or $D_{0}(AB) + P(A, B)$, as appropriate. In comparing the intrinsic bond energies of the TiNb, VNb, and NbCr molecules, one should note that these values form a more understandable pattern than do the directly measured dissociation energies, $D_{0}$. Whereas $D_{0}$ (VNb) is anomalously high compared to both $D_{0}(TiNb)$ and $D_{0}(NbCr)$, the intrinsic bond energies of these molecules show TiNb and VNb to have similar bond energies, while a continuous drop in intrinsic bond energies commences with NbCr. Thus, the chemical bonding is similar in TiNb and VNb, except for the extra promotion energy required to prepare Ti for bonding. The same conclusion is also found when comparing the 3d−3d dimers that are isovalent to these molecules, TiV and V2. In fact, the difference between the intrinsic bond energies of VNb and TiNb, 0.220 eV, almost exactly equals the corresponding difference between V2 and TiV, 0.208 eV.

The large drop in intrinsic bond energy that is observed in going from VNb to NbCr probably also appears in a comparison of V2 to VCr. Unfortunately, our lack of a reliable value of $D_{0}(VCr)$ prohibits a definitive statement on this point. However, a comparison of the intrinsic bond energies of V2 [$D_{0}(V2) = 2.753$ eV] and Cr2 [$D_{0}(Cr2) = 1.53(6)$ eV],4 which are 4.025 and 2.47 eV, respectively, shows a drop of 1.55 eV. This difference compares to a drop in intrinsic bond energy of 0.928 eV when a single V is changed to Cr in moving from VNb to NbCr. Thus, one might expect a similar decrement of 0.77−0.93 eV in moving from V2 to VCr. This allows the VCr bond energy to be predicted to lie in the range of 1.99 eV ≤ $D_{0}$ (VCr) ≤ 2.15 eV. This result compares very well with the bond energy calculated by Andersson, $D_{0}(VCr) = 2.062$ eV.18 The significant drop in intrinsic bond energy that is found in going from VNb to NbCr (0.928 eV), or from V2 to Cr2 (1.55 eV), probably results from the 10% drop in 3d orbital size in going from V to Cr. This allows the VCr bond energy to be predicted to lie in the range of 1.99 eV ≤ $D_{0}$ (VCr) ≤ 2.15 eV. This result compares very well with the bond energy calculated by Andersson, $D_{0}(VCr) = 2.062$ eV.18 The significant drop in intrinsic bond energy that is found in going from NbCr to NbCo and NbNi reflects the decreasing size of the 3d orbitals as one moves across the 3d series. This drop is reflected in all transition metal homonuclear molecule bond lengths, as shown in Table 4.

### Table 4: Bond Energies of NbM and VM Molecules, where M Is a 3d-Series Transition Metal

<table>
<thead>
<tr>
<th>number of d + s electrons</th>
<th>diatomic molecule</th>
<th>$D_{0}$ (eV)</th>
<th>$P(A)$ (eV)</th>
<th>$P(B)$ (eV)</th>
<th>$D_{0} + P_{&lt;}$ (eV)</th>
<th>$D_{0} + P(A) + P(B)$ (eV)</th>
<th>ground state</th>
<th>references</th>
</tr>
</thead>
<tbody>
<tr>
<td>9</td>
<td>TiNb</td>
<td>3.092(1)</td>
<td>1.113</td>
<td>0.524</td>
<td>3.616</td>
<td>4.729(1)</td>
<td>$7\sigma 15d_{2}/2d_{3}, 3\Sigma^{-}$</td>
<td>26, 41</td>
</tr>
<tr>
<td>10</td>
<td>VNb</td>
<td>3.789(1)</td>
<td>0.636</td>
<td>0.524</td>
<td>4.313</td>
<td>4.949(1)</td>
<td>$7\sigma 15d_{2}/2d_{3}, 3\Sigma^{-}$</td>
<td>7, 10</td>
</tr>
<tr>
<td>11</td>
<td>NbCr</td>
<td>3.026(3)</td>
<td>0.524</td>
<td>0.471</td>
<td>3.497</td>
<td>4.021(1)</td>
<td>$7\sigma 15d_{2}/2d_{3}, 3\Delta_{0}$</td>
<td>this work</td>
</tr>
<tr>
<td>14</td>
<td>NbCo</td>
<td>2.729(1)</td>
<td>0.524</td>
<td>0.648</td>
<td>3.253</td>
<td>3.901(1)</td>
<td>$7\sigma 15d_{2}/2d_{3}, 3\Delta_{0}$</td>
<td>25</td>
</tr>
<tr>
<td>15</td>
<td>NbNi</td>
<td>2.780(1)</td>
<td>0.524</td>
<td>0.166</td>
<td>2.946</td>
<td>3.470(1)</td>
<td>$7\sigma 15d_{2}/2d_{3}, 3\Delta_{0}$</td>
<td>25, 42</td>
</tr>
<tr>
<td>9</td>
<td>TiV</td>
<td>2.068(1)</td>
<td>1.113</td>
<td>0.636</td>
<td>2.704</td>
<td>3.817</td>
<td>$5\sigma 15d_{2}/2d_{3}, 3\Sigma^{-}$</td>
<td>20, 33</td>
</tr>
<tr>
<td>10</td>
<td>V2</td>
<td>2.753(1)</td>
<td>0.636</td>
<td>0.636</td>
<td>3.389</td>
<td>4.025</td>
<td>$5\sigma 15d_{2}/2d_{3}, 3\Sigma^{-}$</td>
<td>1, 3, 20</td>
</tr>
<tr>
<td>11</td>
<td>VCr</td>
<td>2.068(1)</td>
<td>0.636</td>
<td>0.471</td>
<td>2.266</td>
<td>2.902(1)</td>
<td>$5\sigma 15d_{2}/2d_{3}, 3\Delta_{0}$</td>
<td>this work</td>
</tr>
<tr>
<td>15</td>
<td>VNi</td>
<td>2.100(1)</td>
<td>0.636</td>
<td>0.166</td>
<td>2.266</td>
<td>2.902(1)</td>
<td>$5\sigma 15d_{2}/2d_{3}, 3\Delta_{0}, 3\Sigma^{+}$</td>
<td>20, 33</td>
</tr>
</tbody>
</table>

Notes:
- Calculated from the atomic energy levels listed in ref 30. The quantities $D_{0} + P_{<}$ and $D_{0} + P(A) + P(B)$ represent the intrinsic bond energies that apply for the cases of a diatomic molecule correlating to a $d_{s1} + d_{s2}$ separated-atom limit and to a $d_{s1} + d_{s2}$ separated-atom limit, respectively.
- Ground states listed in square brackets are not known from experimental measurements or ab initio calculations, but represent reasonable guesses based on the known ground states of similar molecules.

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### V. Conclusion

The ground-state $r_{0}$ bond lengths of VCr and NbCr, 1.7260−(11) and 1.8940(3) Å, respectively, have been determined using rotationally resolved optical spectroscopy. These values are in close agreement with available theoretical predictions for VCr and are close to the averages of the $r_{0}$ bond lengths of the corresponding homonuclear molecules. On the basis of the known $r_{0}$ bond lengths of V2, Cr2, Nb2, Mo2, VNb, CrMo, and now VCr and NbCr, multiple bonding radii are assigned for the intermetallic bonds of the constituent atoms. These radii reproduce the known bond lengths to an accuracy of 0.012 Å and allow the bond lengths of the as-yet-unmeasured species VMo and NbMo to be predicted as 1.861 and 2.016 Å, respectively. The ground state of VCr is assigned as $3\Delta_{0}$, and
the ground state of NbCr is thought to be $2\Delta$, as well, although further experimental work will be required to prove this. The bond energy of NbCr has also been measured to be $24.409 \pm 0.006$ eV by the observation of a sharp predissociation threshold in a dense vibronic spectrum. Finally, the trends in the accurately known bond energies of TiNb, VNb, NbCr, NbCo, and NbNi, along with the isovalent molecules TiV, V$_2$, and VNi, are discussed.

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